Double replacement- A chemical reaction between compounds in which the elements in the reactants recombine to form two different compounds, each of the products having one element from each of the reactants. In the double replacement lab, we combined aqueous Lead Nitrate with aqueous Potassium Iodide to produce solid Lead Iodide and aqueous Potassium Nitrate. We observed two liquids being combined in a beaker, which then immediately produced a yellow solid and liquid remains. In this lab, Potassium switched places with Lead, causing a double replacement to occur.

Pb(NO3)2 (L) + 2KI (L) -> PbI2 + 2K(NO3)

Combustion- A chemical change usually involving oxidation, accompanied by the production of heat and light. In the Combustion lab, we combined Hexane and fire to produce Water and Carbon Dioxide after the fire burned out. The result was the gas produced by Hexane reacting with the air combusted when it reached fire. We then combined Acetone with fire to produce flames that created Carbon Dioxide and Water when it disappeared. In this lab, when a combustible compound hit fire, it produced fire, water and carbon dioxide.

Hexane: C6H14 + O2 -> CO2 + H2O

Acetone: CH3COCH3 +O2 -> Co2 + H2O